

**ATOMIC ENERGY CENTRAL SCHOOL-
KUDANKULAM**

Worksheet –Module-2/4

Subject-Chemistry

Class-X

Lesson No.- Chapter 1 Chemical Reactions and Equations

**Name of the topic – COMBINATION REACTION AND
DECOMPOSITION REACTION**

Total Marks-10x3=30

- (1) Write the chemical equation of the reaction in which the following changes have taken place with an example of each:
 - (i) Change in colour
 - (ii) Change in temperature
 - (iii) Formation of precipitate
- (2) State the type of chemical reactions and chemical equations that take place in the following:
 - (i) Magnesium wire is burnt in air.
 - (ii) Electric current is passed through water.
 - (iii) Ammonia and hydrogen chloride gases are mixed
- (3) Write the balanced chemical equation for the following reaction:
 - (i) Phosphorus burns in presence of chlorine to form phosphorus pentachloride.
 - (ii) Burning of natural gas.
 - (iii) The process of respiration.
- (4) Write one example for each of decomposition reaction carried out in presence of
 - (i) Electricity (ii) Heat (iii) Light
- (5) Which products will be obtained when lead nitrate is heated simply? Write balanced chemical equation for the reaction? State the type of chemical reaction that occurs in the change.
- (6) 2g of ferrous sulphate crystals are heated in a dry boiling tube.
 - (i) List any two observations.

(ii) Name the type of chemical reaction taking place.

(iii) 'Write the chemical equation for the reaction.

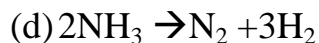
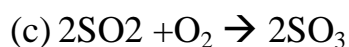
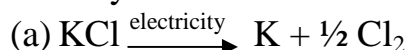
(7) What happens-

(i) Molten alumina is electrolysed

(ii) Silver bromide is kept in sunlight

(iii) Magnesium carbonate is heated

(8) Identify the combination and decomposition reaction in the following-



(9) Give two examples each of exothermic and endothermic reaction.

(10) What is the source of energy for the following decomposition reactions-

(a) Decomposition of potassium chlorate to potassium chloride and oxygen

(b) Electrolysis of brine solution to get hydrogen, chlorine and caustic soda.

(c) Depletion of ozone in stratosphere due to chlorofluorocarbons.

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