**Total No. of printed pages : 2**

SCHOOL NAME:\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
 WORKSHEET – 1/Module-1

Sub: CHEMISTRY Class : XI

Lesson : CLASSIFICATION OF ELEMENTS AND PERIODICITY IN
 PROPERTIES

Name:\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ Roll No.:\_\_\_\_ Date: \_\_\_\_\_\_\_\_\_\_

Maximum Marks : 20 Marks Obtained: \_\_\_\_\_

**I. (MULTIPLE CHOICE QUESTION)MCQS (10X1=10M)**

**1. Zero group was introduced by**

 **a) Lother Meyer b) Mendeleef c) Ramsay d) Lockyer**

**2. The plot of √ν vs Z is**

 **a) straight line b) exponential curve

 c) hyperbolic d) curve with –ve slope**

**3. The statement that is not true for the long form of the periodic table**

 **a) It reflects the sequence of filling the electrons in the order of sub
 energy levels s, p, d and f
 b) It helps to predict the stable valence states of the elements**

 **c) It reflects trends in physical and chemical properties of the elements**

 **d) It helps to product the relative ionic nature of the bond between any
 two elements**

**4. The element with atomic number 19 is**

**a) halogen b) chalcogen c) noble gas d) an alkali metal**

**5. A pair of atomic numbers which belong to s-block are**

 **a) 7,15 b) 6,12 c) 9,17 d) 3,12**

**6. The atomicity of a noble gas is**

 **a)2 b) 1 c) 4 d) 6**

**7. The first lanthanoide is**

 **a) La b)Ce c) Th d) Lu**

**8. The maximum number of valence electrons possible for atoms in the second
 period of the periodic table is**

 **a)18 b)10 c) 8 d) 2**

**9. The atomic number of an element ‘X’ is 34. Then it is present in
 \_\_\_\_\_\_\_period and \_\_\_\_\_\_\_in group.**

 **a) 4th period and 4th group b) 4th period and 6th group**

 **c) 4th period and 7th group d) 5th period and 6th group**

**10. At room temperature liquid metal and liquid non-metals are**

 **a) Hg& I2 b) Cs & Cl2 c) Hg & Br2 d) Cd & S**

 **II. ANSWER THE FOLLOWING QUESTIONS:- (10M)**

1. **Why are lanthanoids and actinoids placed separately at the bottom of the periodic table? (1M)**
2. **Why Zn, Cd, Hg are not considered as transition elements? (1M)**
3. **Why chemistry of actinoids is complicated? (1M)**
4. **Predict block and period of an element whose atomic number is 33.
 (1M)**
5. **The element 119 has not been discovered. What could be the IUPAC name and symbol for this element? And also write the formula of the most stable chloride and hydroxide of that element. (2M)**
6. **A monoatomic anion of unit charge contains 45 neutrons and 36 electrons. What is the atomic mass of the element and in which group of the periodic table does it lie? (2M)**

 **7.On the basis of electronic configurations, Justify that the second
 and third period contains 8 elements each, 4th and 5th periods
 contain 18 elements each and the 6th period contains 32
 elements (2M)
 ….X….**